

Trace element cycling in a subterranean estuary: Part 2. Geochemistry of the pore water

Matthew A. Charette* and Edward R. Sholkovitz
Department of Marine Chemistry and Geochemistry
Woods Hole Oceanographic Institution
Woods Hole, Massachusetts 02543 USA

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*corresponding author: e-mail, mcharette@whoi.edu; tel, 508-289-3205; fax, 508-457-2193

Abstract

Submarine groundwater discharge (SGD) is an important source of dissolved elements to the ocean, yet little is known regarding the chemical reactions that control their flux from sandy coastal aquifers. The net flux of elements from SGD to the coastal ocean is dependent on biogeochemical reactions in the groundwater-seawater mixing zone, recently termed the "subterranean estuary". This paper is the second in a two part series on the biogeochemistry of the Waquoit Bay coastal aquifer/subterranean estuary. The first paper addressed the biogeochemistry of Fe, Mn, P, Ba, U, and Th from the perspective of the sediment composition of cores (Charette et al., 2005). This paper uses pore water data from the subterranean estuary, along with Bay surface water data, to establish a more detailed view into the estuarine chemistry and the chemical diagenesis of Fe, Mn, U, Ba and Sr in coastal aquifers.

Nine high-resolution pore water (groundwater) profiles were collected from the head of the bay during July 2002. There was non-conservative addition of both Ba and Sr in the salinity transition zone of the subterranean estuary. However, the extent of Sr release was significantly less than that of its alkaline earth neighbor Ba. Pore water Ba concentrations approached 3000 nM compared with 25-50 nM in the surface waters of the bay; the pore water Sr-salinity distribution suggests a 26 % elevation in the amount of Sr added to the subterranean estuary. The release of dissolved Ba to the mixing zone of surface estuaries is frequently attributed to an ion-exchange process whereby seawater cations react with Ba from river suspended clay mineral particles at low to intermediate salinity. Results presented here suggest that reductive dissolution of Mn oxides, in conjunction with changes in salinity, may also be an important process in maintaining high concentrations of Ba in the pore water of subterranean estuaries.

In contrast, pore water U was significantly depleted in the subterranean estuary, a result of SGD-driven circulation of seawater through reducing permeable sediments. This finding is supported by surface water concentrations of U in the bay, which were significantly depleted in U compared with adjacent coastal waters. Using a global estimate of SGD, we calculate U removal in subterranean estuaries at $20 \times 10^6 \text{ mol U y}^{-1}$, which is the same order of magnitude as the other major U sinks for the ocean. Our results suggest a need to revisit and reevaluate the oceanic budgets for elements that are likely influenced by SGD-associated processes.

1. INTRODUCTION

The most exciting geochemical processes on the earth occur at interfaces. One major interface that has received limited attention is the groundwater-seawater mixing zone beneath coastlines. The potential importance of this boundary led Moore (1999) to introduce the concept of a “subterranean estuary”, which he defined as the mixing zone between terrestrially-derived fresh groundwater and seawater in a coastal aquifer. Geochemical budgets associated with water transport and reactions in subterranean estuaries are poorly constrained, mainly because they are difficult interfaces from which to collect samples (Burnett et al., 2001, 2002). Recent studies indicate that groundwater-seawater interactions in subterranean estuaries may be important sources of Ba, Ra and Sr to the ocean; the oceanic budgets of these and other inorganic elements need to be reevaluated with respect to processes in subterranean estuaries (Moore, 1996; Shaw et al., 1998; Basu et al., 2001).

A series of papers have argued that the transport of Ba and Ra by submarine groundwater discharge (SGD) can account for a significant fraction of the elevated concentrations of Ba and Ra observed in many estuaries and shelf waters (Moore, 1996, 1997, 1999; Moore and Shaw, 1998; Shaw et al., 1998). These papers suggest that there are two biogeochemical processes responsible for elevating the concentrations of Ba and Ra. The first process involves the release of dissolved Ba and Ra by ion exchange reactions during the landward intrusion of seawater into coastal aquifers. Shaw et al. (1998) and Moore (1999) also suggest that the release of dissolved Ba and Ra to coastal aquifers via SGD may be driven by the redox cycles of iron and manganese oxides. The underlying mechanism in this scenario is the adsorption of Ra and Ba onto the metal oxides under oxic conditions and their release under sub-oxic or anoxic conditions to discharging submarine groundwater.

In a study of strontium in the Ganges-Brahmaputra (G-B) delta, Basu et al. (2001) concluded that SGD was a potential source of strontium to the oceans equal in magnitude to the dissolved Sr concentrations carried to the oceans by the G-B river water. They also suggested that their estimate of SGD-derived Sr was significant enough to have a major effect on the Sr isotope composition of seawater. Hence, the geochemical cycles of three widely-applied alkaline

earth elements (Ba, Ra and Sr) may be influenced by reactions and flow in subterranean regions near estuaries and coasts.

This paper is the second of a two part series in this journal on the biogeochemistry of the Waquoit Bay coastal aquifer/subterranean estuary. The first paper addressed the biogeochemistry of Fe, Mn, P, Ba, U, and Th from the perspective of the sediment composition of cores. (Charette et al., 2005). By focusing on this group of elements, four important classes of reactions (and reactants) are studied. These include [1] redox-controlled solubility (Fe, Mn and U), [2] adsorption (P, Th, Ba, Ra) onto the oxides of Fe and Mn, [3] release (P, Ba, Ra) from oxides undergoing reductive dissolution, and [4] desorption (Ba, Ra) from sediments via ion-exchange reactions. This paper will use ground (or pore) water and Bay surface water data to establish a more detailed view into the estuarine chemistry and the chemical diagenesis of Fe, Mn, U, Ba and Sr in the coastal aquifer. The small and confined nature of the subterranean estuary in Waquoit Bay allows these two papers to provide an in-depth study of the major biogeochemical reactions operating on the permeable sediments of a coastal aquifer with active Fe and Mn redox cycles and well-defined salinity gradients.

“Groundwater” and “pore water” are terms which refer to water within sediments. The term groundwater is usually associated with freshwater systems and has the connotation of flow. Pore water is extensively used in papers on marine systems. These terms will be used interchangeably in this paper on subterranean estuaries.

2. EXPERIMENTAL METHODS AND PROCEDURES

2.1. Field Methods

A series of groundwater (or pore water) profiles were collected from the head of Waquoit Bay during July 2002 (Fig. 1; Table 1). In order to obtain a 2-D representation of the geochemical properties across the aquifer's salinity gradient, we conducted a six station transect perpendicular to the shore line (Fig. 1; line A–A', piezometer stations # 3, 5, 6, 7, 8 and 10). This transect began at the berm of the beach and extended 17-m to a point several meters below the low tide mark. We also sampled at four other stations along a 180 m transect which ran

parallel to the shore (Fig. 1; line B-B' and piezometer stations # 1, 2, 4 and 9). Each profile required 4 to 8 hours of sampling, and the sampling covered 17 days (7 June to 3 July). Hence, the pore water data do not represent synchronous distributions of the measured parameters.

The groundwater samples were obtained with a stainless steel drive point piezometer system ("Retract-A-Tip" from AMS, Inc. Idaho, USA; Allen and Charette, 2004). Water samples were collected at ~0.5 m intervals in the fresh and saline portions of the aquifer and at 0.15 m intervals when the salinity gradient was encountered. Groundwater samples were brought to the surface through acid-cleaned Teflon or polypropylene tubing using a peristaltic pump. After flushing the tubing several times, samples for trace metal analysis were directly filtered into acid cleaned 15 ml LDPE bottles using an all-plastic syringe and a 0.2 μm (pore sized) capsule filter. Each sample bottle was preloaded with 20 μL of trace metal grade nitric acid (8 M). This technique lead to the immediate acidification of each filtered sample to pH 1-2.

We used a YSI 600R multi-probe (Salinity, Temperature, DO, pH) with a flow through cell (YSI, Inc.) to guide our sample collection strategy in the field. In addition, separate discrete water samples were collected for salinity (Guideline AutoSal) and dissolved oxygen (Winkler titration) determinations which were conducted back in the laboratory. A separate filtered sample was also collected and stored frozen for nutrient analysis. Nutrients ($\text{NO}_3^-/\text{NO}_2^-$, PO_4^{3-} , NH_4^+ , SiO_4^-) were analyzed via a flow injection colorimetric method (Lachat QuickChem 8000) and will be discussed in separate manuscript.

Surface water samples were collected from Waquoit Bay during a July 1999 cruise described in Charette et al. (2001), which focused on Ra as a tracer of SGD into the Bay. We analyzed these archived samples, which were collected with a peristaltic pump and syringe filtered ($< 0.2 \mu\text{m}$) into acid cleaned LDPE bottles, for dissolved U and Ba concentrations (Table 2).

2.2. Laboratory Methods and Analysis

The ground (pore) water and surface Bay water concentrations of total dissolved Fe, Mn, Ba, Sr and U were measured by inductively coupled plasma mass spectrometry (ICP-MS) as described in Charette et al. (2005). Briefly, each sample was diluted 1:20 with 1 N nitric acid. Standards were prepared in the same acid. Indium (In), added to samples and standards, served as an internal standard in order to quantitatively account for variations in the performance of the mass spectrometer during analysis. The resulting In count rate data was used to normalize the mass spectrometer count rate for Fe, Mn, Ba, Sr and U. A Finnigan Element High Resolution ICP-MS, operated by the WHOI Mass Spectrometry facility, was used in low-resolution mode for In (115), U (238), and Ba (137) and in the medium-resolution mode for Mn (55), Fe (56), and In (115). Given the high precision and accuracy requirements for the surface water U determination, these samples were analyzed via isotope dilution ICP-MS using ^{232}U as the spike. The analytical methods associated with the sediment compositions are described in Charette et al. (2005).

3. RESULTS and DISCUSSION

The elemental concentrations of the pore water profiles and surface Bay water are contained in tables 1 and 2 respectively. While this paper will only discuss salinity, Fe, Mn, Ba, Sr and U, these tables also contain data for dissolved oxygen, dissolved organic carbon, PO_4^{3-} , and SiO_4^- . Figures 2 and 3 provide two-dimensional distributions of salinity, Fe, Mn, Ba and U. These 2-D plots are derived from the transect perpendicular to the shore line (Fig. 1; line A-A', piezometer stations # 3, 5, 6, 7, 8 and 10). The vertical profiles in figure 4 illustrate the major downward variations in the pore water concentrations of Fe, Mn, Ba, and U from stations PZ7 and PZ3. PZ7 is almost entirely in the freshwater region of the SGD while PZ3 captures the full salinity gradient of the subterranean estuary. Sr data are only available for PZ3.

Figures 5, 6 and 7 compare and contrast the salinity-based distribution of Ba, Sr and U in the pore water and Bay water samples. Figure 5 compares the Ba-salinity distributions of the pore water and Bay water samples. Figures 6a and b, using data from only PZ3, compare the

salinity distribution of Sr directly with that of Ba. Figure 6c is Ba-Sr plot of the same data. Figure 7 compares the U-salinity distributions of the pore water and Bay water samples. It also shows the predicted U-salinity distribution based on conservative mixing between open ocean seawater and the aquifer's freshwater end-member.

We have used Ba-salinity and U-salinity distributions of the surface Bay water in figures 5 and 7 to extrapolate to their zero salinity concentrations. This approach was used by Boyle et al. (1974) as means of estimating the extent to which an element is released to or removed from the mixing zone of estuaries. This type of extrapolation yields what is called the “effective zero salinity end-member concentration”. Coffey et al. (1997) have employed this approach to Ba data from many estuaries. As the salinity range of our Bay water samples is narrow (27.5 to 31.5), our extrapolations for Ba and U carry large uncertainties. As noted later, the extrapolated concentrations of Ba and U are consistent with other evidence for their addition and removal within the subterranean estuary.

Figure 8 plots dissolved Mn against dissolved Ba for pore water samples collected in the middle of A-A' transect of the subterranean estuary at PZ3, PZ6 and PZ8. Finally, because the B-B' transect only includes 4 stations spanning ~200 m versus 6 stations across 17 m for the A-A' transect, we do not include contour plots for these sites. However, the data are presented in Table 1 and discussed throughout the paper.

3.1. Salinity Distribution in the Bay and its Subterranean Estuary

Background into the hydrological and geological setting of the study site has been presented in Charette et al. (2005). A brief overview follows. As shown in Figure 1, Waquoit Bay is a shallow, semi-enclosed and tidally-flushed estuary located on the south shoreline of Cape Cod (MA, U.S.A.). The upper ~10 m of the Cape Cod Aquifer consists of a relatively homogeneous distribution of highly permeable sediment along the shoreline of Waquoit Bay (Cambareri and Eichner, 1998). Because the soil on Cape Cod is primarily composed of coarse-grained sand, precipitation tends to infiltrate the sediments rather than become surface runoff.

Thus, groundwater is the major source of freshwater to the Bay and to the two small rivers that drain into it (Cambareri and Eichner, 1998; Valiela et al., 1990; Charette et al., 2001; U.S. Geological Survey, 2003).

Unlike the piezometer-collected pore water samples, which cover the whole mixing zone (Fig. 2), the salinity of our surface Bay waters only covers a small range (27.6 to 31.5) (Figs. 5-7). Surface water salinity generally increases with distance from the head to the mouth of the bay (Charette et al., 2001). For example, the salinity near the head of the Bay is 27.61 and increases to 31.49 in Vineyard Sound outside the inlet to the Bay. This salinity pattern supports the idea that the head of the bay, the location of our piezometer survey, is an important source of SGD to the system (Charette et al., 2001).

The 2-D distribution of pore water salinity in figure 2a shows that there is a well-defined subterranean estuary beneath the head of the Bay. The main feature is a narrow seepage face where zero salinity groundwater flows upward to the surface of the beach. Seepage meter studies of the intertidal zone confirm that submarine (saline) groundwater discharge also occurs in a narrow (~25 m wide) band along the head of the bay (Michael et al., 2003; Sholkovitz et al. 2003). As Figure 2a reveals, sharp salinity gradients exist over short distances in both the vertical and horizontal directions. For example, salinity increases from 0 to 28 over 2.5 vertical meters in the center of the beach. This same large salinity gradient occurs over a horizontal distance of less than 10 m in the upper meter of beach sediments. These large salinity gradients result in a strong pycnocline, which leads to vertical and horizontal density stratification within the subterranean estuary.

The salinity distribution along the B-B' shore parallel transect (PZ 1-4; Table 1) also shows that there is a sharp boundary between fresh and saline groundwater. The upper sections of profiles from PZ 1, 2 and 4 also show a reverse salinity gradient, a transient features which results from tidally pumped seawater into the upper ~1 m of the sediment column during high tides.

3.2. REDOX Framework of the Subterranean Estuary: Fe and Mn

This section sets the REDOX (oxidation-reduction) framework of the subterranean estuary, which is described in more detail in both Charette and Sholkovitz (2002) and Charette et al. (2005). The 2-D distributions of pore water Fe and Mn along transect A-A' and the vertical profiles from PZ7 and PZ3 (Figs. 2 and 4) show that the subterranean estuary in Waquoit Bay has two distinct sources of the high concentrations of total dissolved and reduced Fe and Mn. One source resides in the upward rising plume of freshwater, and the second source lies in the salt-wedge zone of mid to high salinity pore water. The distribution of Fe and Mn within the *freshwater* region of the subterranean estuary shows that there is a plume of Fe and Mn rich groundwater moving toward the Bay (Figs. 2 and 4). This plume is centered between a depth of 2 and 4 m and carries up to 175 μM of Fe and 15 μM of Mn. As the leading edge of this plume moves upward toward the beach, almost all of the Fe and Mn are removed over a horizontal and vertical distances of only a few meters. (Our measurements of ferrous and total dissolved iron concentrations by the ferrozine colorimetric method of Stookey (1970) show that the major fraction of the total dissolved Fe in this subterranean estuary was in the reduced or ferrous form).

The second source of dissolved Fe and Mn lies in the salt wedge section of the subterranean estuary where pore water concentrations reach 50 μM of Fe and 40 μM of Mn (Figs. 2 and 4). These high concentrations of Fe and Mn are associated with a pore water salinity that exceeds 5-10. Another key redox feature in the salt-wedge section of the subterranean estuary is the fact that a lens of high Mn concentration lies over a lens of high Fe concentration (Figs. 2 and 4)

Unlike the A-A' transect, the stations from the shore parallel (B-B') transect were too far seaward to capture the Fe and Mn plumes of the zero salinity groundwater. However, the B-B' distribution of Fe and Mn in the salt-wedge section of the subterranean estuary were similar to the A-A' transect in that the Mn concentration peaks were generally above the Fe concentration maxima. The concentration ranges were much higher than those of the A-A' transect. Mn approaches 100 μM at a depth of $\sim 2\text{m}$ in PZ1 and Fe reaches a peak of nearly 500 μM at $\sim 8\text{ m}$ in PZ4.

Charette and Sholkovitz (2002) collected ~ 2 m long sediment cores from the fully saline and intertidal section of the subterranean estuary at the head of Waquoit Bay (near PZ8 in Fig. 1). The most outstanding visual features of these cores are their colors. Charette and Sholkovitz (2002) showed that the dark red, yellow and orange colors of the Waquoit Bay cores were iron (hydr)oxide precipitates loosely bound to sand grains. The deeper sections of the cores are characterized by large amounts of iron (hydr)oxides (ferrihydrite, lepidocrocite and goethite) that are being precipitated onto organic C-poor quartz sand (Charette et al. 2005). Unlike Fe (hydr)oxides which increase with depth, the Mn (hydr)oxides display well-developed maxima in the mid-sections of cores 2 and 3. This type of vertical stratification is consistent with redox-controlled diagenesis in which Mn (hydr)oxides are formed at shallower depths than iron (hydr)oxides. The reduction of seawater-derived sulfate and the production of sulfides are not important parts of the redox-driven cycles in this system. As noted later in the context of their pore water gradients, the sediment profiles are also characterized by enrichments in Ba and U.

In summary of this section, all of the observations point to active Fe and Mn redox cycles in this coastal aquifer, even though the concentration of solid-phase organic carbon in the Waquoit Bay sediments is low (< 0.1 %; Charette et al., 2005). Groundwater from the land delivers high concentrations of dissolved ferrous iron and reduced manganese to the freshwater end of the subterranean estuary where a large fraction of this dissolved Fe and Mn is oxidized and precipitated onto the quartz beach sands. There is also production, transport and subsequent oxidation of reduced Fe and Mn at the high salinity end of the subterranean estuary where the formation of iron (hydr)oxide rich sands are forming (Charette and Sholkovitz, 2002; Charette et al., 2005).

3.3. Barium in the Subterranean Estuary and the Bay

A major feature of this study is the large-scale enrichment of Ba in the subterranean estuary of Waquoit Bay with respect to the concentrations of Ba in both the freshwater and seawater end-members (Figs. 3-5). Barium concentrations span nearly four orders of magnitude from ~1 nM to 2800 nM (Figs. 3-5). Coastal seawater in Vineyard Sound contains ~25 nM of dissolved Ba; the freshwater portion of the aquifer is also characterized by low Ba concentration (10-50 nM). As shown in the 2-D distributions of the shore-perpendicular transect (Fig. 3), Ba

reaches high concentrations at the upper side of the salinity transition zone (STZ). This feature is most pronounced in PZ3 where Ba concentrations increase from 66 nM at a depth of 1.52 m and a salinity of 0.46 to a value of 1227 nM at a depth of 2.29 m and a salinity of 15.1. The concentration of Ba decreases with increasing depth and salinity. Hence, Ba exhibits a subsurface maximum that spans ~ 1.5 m and lies between the zones of fresh and saline pore water.

The Ba-salinity distribution for all the pore water samples confirms that Ba follows a strongly non-conservative mixing distribution (Fig. 5a). Again, the Ba concentration is highest in the STZ and generally decreases from mid salinity (~20) to high salinity. The low Ba values in a handful of STZ samples are associated with shallow samples from the piezometer profiles, and are, therefore, an indicator of recent seawater infiltration. In B-B' transect along the shoreline there is a clear trend of decreasing Ba from west (PZ1) to east (PZ4) (Table 1). Maximum Ba concentrations decreased from 2800 to 260 nM.

Given the scatter when plotting all of the pore water Ba against salinity (Fig. 5a), it is difficult to estimate the “effective” zero-salinity concentration within the subterranean estuary. We have done so using the Ba-concentration diagram for PZ3 only (Fig. 6b). When one extrapolates from the high salinity zone back to zero salinity, the ‘effective’ Ba concentration is ~ 2700 nM. In context, the measured Ba concentration of the fresh ground water entering the subterranean estuary is 10-50 nM, or two orders of magnitude lower (Fig. 3).

As shown in Figures 2-4, a prominent feature of our pore water data set is that the subsurface maximum in Ba coincides with a lens of elevated Mn concentrations. The lens of high Ba and Mn are both associated with the STZ. In contrast, the lens of high dissolved Fe concentration lies much deeper in the salt wedge region of the subterranean estuary. As will be discussed later, our data suggest that the redox geochemistry of Mn (and not Fe) has an important controlling role in forming the elevated concentrations of Ba.

Within the surface Bay water samples, Ba concentrations increase with decreasing salinity (Fig. 5b). Barium in the high salinity waters of Vineyard Sound is ~25 nM and increases

to 55-75 nM at the lowest salinity (27.6-28.0) samples near the head of the Bay. Though the Ba pattern within the Bay are somewhat scattered, highest values were generally associated with stations at the head of the Bay. If one extrapolates the Ba concentration back to zero salinity, then the “effective” Ba concentration is approximately 250 nM (Fig. 5b). Hence, the ‘effective’ freshwater end-member concentration of Ba for the surface samples is 5-25 times higher than the Ba concentration of the ground water and sea water flowing into the subterranean estuary. The “effective” Ba concentration for the pore water samples of PZ3 (Fig 6c; 2700 nM) is ~ 11 times higher than its counterpart for the Bay’s surface waters (250 nM)

3.4. Strontium in Subterranean Estuary and the Bay

Strontium was not part of our initial suite of trace metals analyzed. To date we have only analyzed samples from PZ3 (Figs. 4 and 6). Groundwater concentrations ranged from 0.2 μM at a salinity of ~0 to 84 μM at a salinity of 27. Strontium was only measured on three surface samples from the Bay where the Sr concentrations range from 68-76 μM and are positively correlated with salinity. The three samples fall along a mixing line between the fresh groundwater end-member and the mean ocean water concentration of Sr (Pilson, 1998). Hence, the Sr-salinity distribution indicate that Sr is conservative within the Bay itself, albeit with only three samples as evidence.

The Sr-salinity distribution in figure 6a shows that dissolved Sr behaves non-conservatively within the subterranean estuary. The extent to which Sr is released to the subterranean estuary depends on the choice of the high salinity end-member. If we use the linear mixing line between local coastal sea water, Bay surface water and zero salinity groundwater in Figure 6a as the conservative distribution, then the Sr-salinity distribution suggests a 26 % elevation in the amount of Sr added to the subterranean estuary. This value was calculated by measuring the area under pore water Ba-salinity distribution relative to the area under the conservative mixing line in Figure 6a.

We have also estimated the “effective zero-salinity Sr concentration” of pore water samples from PZ3 by extrapolating the high salinity distributions back to zero salinity (Fig. 6a). This yields an effective end-member concentration of ~ 20 μM . Unlike Ba, the effective

concentration of Sr is much less than its seawater end-member concentration (~ 80 nM). But it is significantly higher than the measured fresh water end-member (~ 0.1 nM).

3.5. A Comparison of Ba and Sr Distributions in the Subterranean Estuary and Bay

A comparison of Ba and Sr in both the pore water and surface waters of Waquoit Bay shows that these two alkaline earth elements differ greatly in their salinity distribution and extent of their estuarine geochemistry (Fig. 6). While both elements are released to the pore water of the subterranean estuary, the extent of release is much greater for Ba than Sr (Fig. 6b). Using the PZ3 data, this distinct difference is shown by comparing their depth profiles (Fig. 4) and by plotting the data as a Ba-Sr scatter diagram (Fig. 6c). Ba and Sr both increase with depth during the transition from fresh water to intermediate salinity. The profiles then diverge. The Sr concentration continues to increase slightly with depth and with increasing salinity; the Sr concentration at the bottom of the profiles is similar to that of the surface water at the same salinity. Ba, in marked contrast, decreases from a high of 1200 nM to values of 140 nM at the bottom of the profile; Bay water at the same salinity is ~ 75 nM. This divergence in the two profiles reflects fundamental differences in their chemical properties.

The surface water of the Bay also shows different Sr-salinity and Ba-salinity distributions (Fig. 5b and 6a). Ba increases, while Sr decreases, with decreasing salinity. Sr appears to be conservative, while Ba, as noted in the last section, is non-conservative. These data suggest that the production of high Ba concentrations within the subterranean estuary are responsible for the excess concentration of dissolved Ba in the overlying water column of the Bay over that expected from conservative mixing. In contrast, the release of Sr to the subterranean estuary is not large enough to measurably affect the Sr-salinity distribution of the Bay water.

3.6. Uranium in the Bay and Subterranean Estuary

Another major feature of this study is that U is strongly non-conservative in this subterranean estuary with large-scale removal over the entire salinity range, and some evidence for U release at high salinity end (Fig. 7a). The sea water concentration outside the Bay is ~ 12 nM, and the U in the freshwater plume ranges from below detection to 2.6 nM. It is clear from

the U-salinity distribution that the removal of sea water- derived U in STZ of the subterranean estuary is nearly quantitative.

The pore water concentration of U in subterranean estuary ranges from below detection to ~20 nM (Figs. 3, 4 and 7). The range in U concentrations for the STZ is similarly low. The high salinity groundwater in the salt-wedge range from detection limit to ~20 nM; the majority of the high U values are located in a narrow (25-26) salinity range. Above this salinity, the U concentration range was much smaller (~0-4 nM). Most of the elevated U in the subterranean estuary are located in the high salinity samples from the middle piezometers (PZ6, PZ3) of the shore-perpendicular transect (Figs. 3 and 4).

Uranium concentrations of the Bay's surface water lie within a narrow range (9.7 – 12.2 nM) and are positively correlated with salinity (Fig. 7a; $R^2 = 0.80$). However, the U-salinity distributions fall below that of the mixing line that connects the coastal sea water of Vineyard Sound to the zero-salinity groundwater entering the Bay (Fig. 7b). Extrapolation to zero salinity yields an “effective” U concentration of -5 nM, which implies that U is being removed from the Bay's surface water. It appears that our finding of U-depleted surface waters supports our hypothesis of large scale U removal within the subterranean estuary.

4. BIOGEOCHEMICAL PROCESSES IN THE SUBTERRANEAN ESTUARY

4.1. Barium

This study provides unequivocal evidence that biogeochemical reactions can lead to the large-scale enrichment of Ba in subterranean estuaries. To the best of our knowledge, Windom and Niencheski (2003) provide the only other known Ba data across a salinity gradient in a coastal (Brazilian) aquifer. Ba concentration in their groundwater profile peaked at 2000 nM at a salinity of ~12. These sets of observations are consistent with recent studies suggesting that groundwater is a major source of Ba to the surface waters of estuaries and the coastal ocean (Moore, 1997; Shaw et al., 1998). What processes control Ba cycling in subterranean estuaries, and hence the flux of Ba to the ocean via SGD? In this section we discuss the following possibilities: (1) ion exchange reactions with seawater, and (2) the redox cycle of Mn, and (3)

weathering of Ba-bearing minerals. We will begin with a brief overview of and comparison with “surface estuaries” where rivers are the primary source of fresh water.

The non-conservative behavior of dissolved Ba is a universal and long-established feature of surface estuaries. Ion exchange, whereby seawater cations substitute for Ba on the surface of river-borne suspended minerals, has long been identified as the most important mechanism for the release of Ba to surface estuaries. Coffey et al. (1997) summarize the behavior of Ba in a wide range of surface estuaries. Large river estuarine systems are often characterized by broad Ba maxima in the low (0-6) salinity region. In contrast, estuaries of small rivers tend to have their Ba concentration maxima at a higher salinity range (10-15). Coffey et al. (1997) argue that the transit time of Ba-releasing particles through an estuary will be the major factor in controlling the salinity region over which a Ba maxima is developed. Coffey et al. also note that the desorption of Ba off river suspended matter occurs on the time scale of minutes to hours at low salinity ($> \sim 2$). The non-conservative distribution of Ba in surface estuaries has been successfully modeled by invoking ion-exchange reactions between sea water and fresh river particles (Hanor and Chan, 1979).

The principal source of dissolved Ba to the mixing zone of a surface estuary is the delivery of river suspended particles (Coffey et. al., 1997). This leads to a continual source of “new” Ba that can be released from fresh river particles at low to intermediate salinities. While some new fine-grained particles may be transported by ground water and recirculating sea water, subterranean estuaries like Waquoit Bay are not supplied with a continual source of fresh mineral particles. The bulk of the sediments in most coastal aquifers are stationary, “old” and more coarse-grained than the suspended particles that characterize rivers and surface estuaries.

When the data for all the pore water samples are used in a Ba-salinity scatter diagram (Fig. 5), one feature of note is that the maximum in the concentration of Ba is skewed toward the higher (15-25) end of the salinity gradient (Fig. 5). Hence, the addition of dissolved Ba to this subterranean estuary appears to occur at salinities which are higher than those associated with most surface estuaries (Li and Chan, 1979; Coffey et al., 1997). The Ba-salinity distribution for PZ3 profile is the only exception to this observation as the peak in Ba is centered on a salinity of

15; its overall distribution is similar to many surface estuaries (Fig. 6b). Thus, the mechanisms driving Ba enrichment in surface estuaries cannot completely account for the Ba excess observed in all subterranean estuaries, except in the case of constant landward intrusion of seawater into coastal aquifers (Moore, 1996).

While Ba and the other alkaline earth elements (Sr, Ra) do not have a redox chemistry of their own, we argue that the elevated Ba concentrations in the pore water of Waquoit Bay are caused in part by the reduction and oxidation cycles of Mn. It is well known that Ba (and Ra) has an exceptionally strong affinity for Mn (hydr)oxides (Balistreri and Murray, 1986; Sugiyama et al., 1992; Mishra and Tiwary, 1993). In fact, MnO₂ impregnated acrylic fibers are used to quantitatively extract Ra from seawater (Moore and Reid, 1976). We suggest a process that is very similar to the coupled redox cycle of Mn and Ba that De Lange et al. (1990) applied to the interface between oxic seawater and anoxic brine water in Bannock basin. Their basic idea is that Ba is recycled at the oxic/anoxic boundary where high concentrations of Ba are maintained as part of the redox cycle of Mn. In this scenario, Ba is adsorbed onto Mn oxides formed from the oxidation of dissolved Mn(II) and released back to the deeper water column as reducing conditions dissolve the Mn oxides. Our field evidence for both the pore water and sediments (Charette et al., 2005) support a coupled redox and salt cycle for Ba. Indeed, the 2-D distributions across transect A-A' and the vertical profiles in PZ3 show that the lens of high pore water Ba is coincident with both a lens of high dissolved Mn (Figs. 2-4) and a strong gradient in salinity. Mn-Ba concentration plots show a strong inter-element correlation for pore water samples; this is not the case for Fe-Ba plots (Fig. 8).

How can this Mn-redox driven cycle of Ba operate in a system that is dominated by the Fe-redox cycle? The answer lies in the fact that the affinity of Ba for Mn oxides is several orders of magnitude greater than the affinity of Ba for Fe oxides (Balistreri and Murray, 1986; Sugiyama et al., 1992; Mishra and Tiwary, 1993; Tonkin et al., 2004). This difference is consistent with the fact that the lens of high pore water Ba in subterranean estuary lies coincident with a lens of Mn but not Fe (Figs. 2-4). The redox cycle of Fe leads to the production of dissolved Fe at a depth underlying that of Mn and Ba.

The Ba-Mn dynamics in the subterranean estuary of Waquoit Bay are likely to be reversible under certain conditions. That is, dissolved Ba will most likely be co-precipitated upon the oxidation of dissolved Mn to Mn (hydr)oxides. Thus, it is possible that there is Ba *removal* at the low to mid salinities that would otherwise be characterized as a zone of large Ba enrichment. Coffey et al. (1997) note that Ba removal occurs in the low salinity zones of several surface estuaries; they suggest removal by oxide formation. In this scenario, a subterranean estuary would be characterized by a coupled Ba-Mn co-precipitation cycle in the low salinity region that maintains Ba concentrations at low (but still highly enriched over surface waters) concentrations while the peak at higher salinity is controlled by Mn (hydr)oxide reduction as described above. The profiles in PZ3 point to the removal of dissolved Ba in association with the removal of dissolved Mn by oxidation to Mn oxides (Fig. 4).

A third factor that could explain the distribution of Ba in the subterranean estuary is the slow weathering of aquifer solids. In contrast with the surface estuary, which continually receives “new” Ba through the transport of suspended particles, the particles that comprise a subterranean estuary are essentially static. Thus, in a scenario where the sea level and aquifer recharge is relatively constant (hence, a relatively constant location of the groundwater-seawater mixing zone on decadal time scales), one would expect the desorptive input of Ba (and other alkaline earth elements) to be minimized. Only through the exposure of newly ion exchangeable Ba via slow weathering of aquifer minerals could this non-conservative release of Ba be maintained. Since silicate has been shown to be a good indicator of mineral weathering in groundwater systems, one piece of evidence in favor of the slow weathering hypothesis is the highly enriched silicate concentrations we observed in the Waquoit Bay subterranean estuary (Table 1). For example, Rogers and Bennett (2004) noted that microorganisms, both directly and indirectly, can lead to the dissolution of silicate bearing minerals in groundwater systems.

Another interesting aspect of the Ba cycle in the Waquoit Bay subterranean estuary is that it is supersaturated with respect to barium sulfate (barite) formation at nearly all depths and locations (Fig. 5). Charette et al. (2005) reported that sulfate concentrations were not depleted to any significant extent in pore water samples from PZ3. Therefore, non-conservative sulfate behavior is not a likely explanation for the high Ba values we observed. This leaves two likely

scenarios: 1) the rate of Ba flux to surface waters exceeds the rate of barite formation in the subterranean estuary and/or 2) the excess Ba is complexed by DOC.

Ba supersaturation has been observed in many other coastal groundwater settings (Shaw et al., 1998; Windom and Niencheski, 2003; Duncan and Shaw, 2003). The relatively fast groundwater flow rates in Waquoit Bay (Cambareri and Eichner, 1998; Charette et al., 2001; Abraham et al., 2003) favor the hypothesis that barite formation kinetics are the major control on Ba concentration in our study. Laboratory experiments of barite precipitation in decaying plankton microenvironments suggest that the process occurs on a time-scale of weeks to several months (Ganeshram et al., 2003), significantly longer than the residence time of groundwater in the subterranean estuary of Waquoit Bay (days to weeks). Conversely, DOC concentrations are sufficiently high (50-700 μM) in this setting that Ba-DOC complexation may also play a factor as has been suggested in many studies of REE in aquatic environments (Duncan and Shaw, 2003).

4.2. Biogeochemistry of Sr in the Subterranean Estuary

The sorption affinity of alkaline earth elements to Mn oxides has the following order, $\text{Ra} \gg \text{Ba} \gg \text{Sr}$ (Balistrieri and Murray, 1986; Sugiyama et al., 1992; Tonkin et al., 2004;). Hence, the release and removal of Sr in conjunction with a Mn-redox cycle is expected to be much less strong than that of Ba. This prediction is consistent with our observation for the subterranean estuary of Waquoit Bay where Sr is much more conservative than Ba (Figs. 4 and 6). The excess inventory of Sr within this subterranean estuary represents, at most, a 26 % elevation over that expected from conservative mixing (Fig. 6a). Studies of seasonally anoxic lakes also show that Ba is much more biogeochemically reactive than Sr (Sholkovitz, 1985; McGrath et al., 1989). Anoxic freshwaters have elevated concentrations of Ba but not Sr.

In a study of Sr and Sr isotopes in the Ganges-Brahmaputra (G-B) delta, Basu et al. (2001) concluded that “subsurface flowing groundwater’s ...is a potential source of strontium to the oceans, equal in magnitude to the dissolved strontium concentrations carried to the oceans by the G-B river water” and have “...an enormous effect on interpreting the seawater Sr isotope balance.” However, these investigators looked only at Sr in the freshwater portion of the aquifer

and did not take into account the potential for Sr and Sr isotopes to be modified during transit through the subterranean estuary. Our Sr results indicate that submarine groundwater discharge could be even more important in controlling the marine Sr budget than Basu et al. (2001) realized.

In their estimate of the global flux of Sr due to SGD, they used a groundwater end member of $\sim 7 \mu\text{M}$. Our data indicate that this value could be much higher ($\sim 20 \mu\text{M}$; Fig. 8) if Sr proves to be non-conservative in other coastal aquifers. This result has significant implications for the Sr isotopic record on geologic time-scales. In a ~ 20 year study of Ra isotopes in a coastal salt marsh, Charette et al. (2003) noted that seasonal/interannual changes in the water table elevation of coastal aquifers are a major factor in alkaline earth elemental fluxes to the ocean through SGD. Similarly, salt water intrusion into coastal sediments during periods of sea level rise may have resulted in significantly elevated inputs of elements like Sr, Ba, and Ra. Indeed, Cochran et al. (2003) invoked the subterranean estuary concept to explain the Sr isotopic record of fossils from the Late Cretaceous Western Interior Seaway. These lines of evidence are supporting SGD as a major, yet unconstrained, non-steady state source of alkaline earth elements to the ocean that may be so significant as to change our interpretation of certain paleoproxies that rely on Ba or Sr/Sr isotopes.

4.3. Geochemistry of Uranium in the Subterranean Estuary

In contrast with the alkaline earth elements, the Waquoit Bay subterranean estuary is a sink for uranium; our U-salinity distribution shows strong evidence of removal at intermediate salinity and some evidence of U release near bay water salinity (Fig. 7). Uranium is typically conservative in oxic seawater where its oxidation state is +6 and it forms a strong complex with the carbonate ion (Koide and Goldberg, 1963). One exception would be the presence of Fe (hydr)oxides, for which U has a relatively strong affinity, especially in freshwater. Under reducing conditions such as in organic rich sediments or oxygen poor ground water, U(IV) is very particle reactive and therefore reducing environments represent a sink for U (Coward, 1980; Cochran et al. 1986). Which of these U removal mechanisms is most important in the Waquoit Bay subterranean estuary?

There exist many case studies of uranium displaying both conservative and non-conservative mixing behavior in surface estuaries. Several studies have shown the Amazon River estuary to be a net sink of U. Removal processes at low to intermediate salinities include colloidal aggregation, flocculation, and adsorption onto resuspended sediments (e.g. Swarzenski et al., 1995; Swarzenski and McKee, 1999). Swarzenski et al. (2004) observed nearly conservative mixing of U in the Fly River estuary (Paupa New Guinea), though U-isotopes indicated that the element is being actively cycled between the water column and sediments. In contrast, Shaw et al. (1994) found that the U distribution in the Chesapeake Bay estuary varied with the redox state of the water column; uranium is removed to the sediments during summertime anoxia and then released back to the water column during other (oxic) parts of the year. Despite these case studies, Windom et al. (2000) noted that U removal in estuaries is the exception rather than the rule, except in the case where a salt marsh environment is present.

Two studies comparable to ours also indicate significant removal of seawater-derived U in coastal aquifers. Both Windom and Niencheski (2003) and Duncan and Shaw (2003) observed significant U depletion at intermediate salinities, which they attributed to removal under anoxic conditions. Charette and Sholkovitz (2002) reported on the formation of “iron curtain” type sediments in the subterranean estuary beneath Waquoit Bay. Subsequent analysis of the U content in the Fe-enriched cores indicates that U is not enriched in the Fe oxide-forming region but rather in the shallow reducing sections of the cores (Charette et al. 2005). It appears that the presence of Fe (hydr)oxides in large amounts is not important in controlling U removal in this environment. It is more likely that the reduction of U(VI) to U(IV) is controlling U removal in Waquoit Bay and in other subterranean estuaries (e.g. Windom and Niencheski, 2003).

Seven of our high (20-30) salinity groundwater samples have U concentrations that exceed the predicted concentrations based on the mixing line with Bay water as the high salinity end member (Fig. 7). Though there is no strong correlation between dissolved oxygen and U, we suspect that the high U values are due to the release of adsorbed U(IV) under more oxidizing conditions. Shaw et. al. (1994) observed a seasonal pattern of removal and release in the estuarine waters of Chesapeake Bay. Removal of U to the sediment occurs during the summertime periods of bottom water anoxia; release to the water column from the sediment

occurs during oxic conditions. Cochran et al. (1986) found that the exposure of coastal marine sediments to the air led to a release of U in excess of the local seawater concentration. Barnes and Cochran (1990) demonstrated that the return of oxic conditions in the sediments U release was rapid (minutes to hours) and quantitative. Hence, one would expect large-scale release of U in our subterranean estuary upon the return of oxygenated pore water to the mixing zone.

The large-scale U removal observed in the subterranean estuary is reflected in the U-salinity distribution of the surface water of Waquoit Bay (Fig. 7). We believe that this is the first water column evidence of U removal in the coastal ocean that can be unequivocally attributed to SGD. Whereas samples from outside the bay followed the U-salinity relationship of Chen et al. (1986), all samples within the bay fell below the line. Extrapolating the Waquoit Bay U trend back to the zero salinity intercept yields an effective groundwater end member of about -5 nM, or about a 50% removal efficiency on average (Fig. 7). Using a Waquoit Bay SGD flux of 30 m y^{-1} as determined by Abraham et al. (2003), we estimate that our subterranean estuary removes $150 \mu\text{mol U m}^{-2} \text{ y}^{-1}$ (Table 3). Our estimate is slightly higher than that of Duncan and Shaw (2003) who estimated an SGD-derived removal flux of $30\text{-}60 \mu\text{mol U m}^{-2} \text{ y}^{-1}$. Our In contrast, Windom et al. (2000) estimated that salt marshes remove $70 \mu\text{mol U m}^{-2} \text{ y}^{-1}$ and Cochran (1982) reported the marine sediments flux to be $0.4 \mu\text{mol U m}^{-2} \text{ y}^{-1}$.

How important might SGD-derived removal of U be on a global basis? With global estimates of SGD ranging from $4,000$ to $100,000 \text{ km}^3 \text{ y}^{-1}$ (Burnett et al. 2003), and a groundwater end member as observed for the Waquoit Bay system of -5 nM, U removal in subterranean estuaries could range anywhere from $20\text{-}500 \times 10^6 \text{ mol U y}^{-1}$ (Table 3). Clearly, given that total oceanic U inputs are not likely much greater than $60 \times 10^6 \text{ mol U y}^{-1}$, our upper-limit estimate is unrealistic. However, such a removal flux would put our lower-limit SGD-derived U sink on par with the global U sink for marine sediments and salt marshes (both $\sim 27 \text{ mol } 10^6 \text{ U y}^{-1}$). These results are in contrast to the recent work of Dunk et al. (2002), who invoke SGD as a net source of U to the oceans ($9.7 \times 10^6 \text{ mol U y}^{-1}$) to achieve a balance for the marine U budget. Even if some coastal aquifer systems are a net source of U to the oceans (e.g. in Karst-type aquifers), this flux is likely offset or even exceeded by removal of U via seawater circulation through coastal permeable sediments as described here.

5. CONCLUSIONS

This study provides unequivocal evidence that biogeochemical reactions can lead to the large-scale enrichment of Ba and depletion of U in subterranean estuaries. This observation is consistent with recent studies suggesting that groundwater is a major source of Ba to estuaries and the coastal ocean (Moore, 1997; Shaw et al., 1998). In addition, the surface water distributions of Ba and U in Waquoit Bay reflect the input and loss processes operating in the subterranean estuary.

A comparison of Ba and Sr distribution in both the pore water and surface waters of Waquoit Bay shows that these two alkaline earth elements differ greatly in their salinity-based distributions and the extent of their estuarine geochemistry. While both elements are released to the pore water of the subterranean estuary, the extent of release is much greater for Ba than Sr. Ion exchange, whereby sea water cations substitute for Ba on the surface of river-borne suspended minerals, has long been identified as the most important mechanism for the release of Ba to surface estuaries. Here, we argue that, in addition to salinity effects, the cycling of Ba in subterranean estuaries is caused at least in part by the redox cycles of Mn. The sorption affinity of alkaline earth elements to Mn oxides has the following order, $Ra \gg Ba \gg Sr$. Hence, the release and removal of Ba, in conjunction with a Mn-redox cycle, is expected to be much stronger than that of Sr. This prediction is consistent with our observation for the subterranean estuary of Waquoit Bay where Sr is much more conservative than Ba. In this context, one would predict that the subterranean estuarine geochemistry of Ra, a commonly applied tracer of SGD, would be more Ba-like than Sr-like.

The redox-driven removal of U in the subterranean estuary is reflected in the surface water U-salinity distribution of Waquoit Bay. We believe that this is the first evidence of U removal in the coastal ocean that can be unequivocally attributed to SGD. Whereas samples from outside the bay followed the U-salinity relationship of Chen et al. (1986), all samples within the bay fell below the line. If the removal of U in our coastal aquifer is scaled to the global estimates of SGD, then our lower-limit of SGD-derived U sink is on par with estimates of

the global U sink for marine sediments and salt marshes. These results are in contrast to the recent work of Dunk et al. (2002), who invoke SGD as a net source of U to the oceans to achieve a balance for the marine U budget.

It remains to be seen if the extent of the observed elemental source/sink mechanisms discussed here are persistent year round and consistent from year to year. For example, seasonal/interannual changes in the water table elevation of coastal aquifers may be a major factor in alkaline earth elemental fluxes to the ocean through submarine groundwater discharge. Similarly, salt water intrusion into coastal sediments during periods of sea level rise may have resulted in significantly elevated inputs of these elements. This could have serious implications for the interpretation of marine paleoproxies such as Sr isotopes (e.g. Cochran et al. 2003). Our Sr results indicate that submarine groundwater discharge could be even more important in controlling the marine Sr budget than Basu et al. (2001) realized. Hence, both present day and long (geologic) time-scale budgets for these and other elements with similar chemical properties (e.g. Mo, Re) must consider the submarine groundwater discharge-associated processes discussed here.

Because of the magnitude and range of advective flow and exchange in coastal aquifers, chemical gradients alone are not sufficient predictors of flux to the overlying water column. Thus, a key component of future research into the subterranean estuary would be a combined geochemical and hydrological flow model. Such a model would improve the applicability of our results to other settings by allowing for the evaluation of new hypotheses based on more complex groundwater-seawater mixing zones.

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Table 1. Pore water concentrations of trace metals and ancillary water quality parameters from the July 2002 piezometer survey.

Station (PZ) locations are shown in Figure 1.

Depth (m)	Salinity	Diss. O ₂ (mg L ⁻¹)	PO ₄ ³⁻ (μmol L ⁻¹)	SiO ₄ ²⁻ (μmol L ⁻¹)	Fe (μmol L ⁻¹)	Mn (μmol L ⁻¹)	Ba (nmol L ⁻¹)	Sr (μmol L ⁻¹)	U (nmol L ⁻¹)	DOC (μmol L ⁻¹)
PZ 1										
0.46	4.80	4.83	0.01	140	0.6	1.5	27		0.1	
0.91	7.47	4.77	0.01	153	0.8	0.8	30		0.4	
1.07	7.85	5.37	0.01	131	0.3	1.4	23		0.4	
1.22	7.07	5.51	0.01	127	0.3	1.0	17		0.1	
1.52	0.63	6.21	0.07	48.6	0.5	0.3	4		0.0	
1.83	0.07	6.26	0.05	30.5	0.1	0.4	1		0.0	
2.16	3.23	0.40	1.55	192	1.2	132.5	2782		0.1	
2.29	18.84	0.39	3.28	320	1.2	96.2	2229		0.1	
2.44	24.39	0.45	2.85	275	1.0	51.4	1543		0.3	
2.59	25.14	0.46	2.78	198	0.8	44.6	652		3.1	
2.74	25.64	0.53	3.2	171	0.7	49.8	386		8.0	
2.90	25.96	0.53	3.92	161	2.2	57.7	355		3.8	
3.35	26.77	0.46	6.32	142	2.0	63.9	384		3.4	
3.51	27.01	1.27	6.03	141	0.7	88.5	961		0.0	
3.96	24.17	3.69	0.48	247	1.1	77.1	509		4.8	
4.42	25.36	1.92	2.5	203	1.1	80.6	424		5.6	
4.88	26.60	2.44	4.22	175	1.5	87.0	467		3.9	
5.33	27.27	2.20	5.22	157	1.3	93.2	480		2.5	
5.94	27.34	1.77	5.74	131	0.7	98.3	667		2.2	
6.55	27.58	2.02	4.18	138	0.6	90.7	682		3.0	
7.01	27.55	2.08	8.25	129	44.1	52.7	298		2.7	
7.32	27.68	6.31	3.8	131	31.1	45.9	282		2.6	
7.47	27.71	1.65	9.2	125	60.0	41.7	223		2.4	
PZ 2										
0.00	25.21	6.09	0.7	18.2	69.7	1.6	61		10.0	
0.46	0.06		2.8	29.9	1.8	0.4	8		0.5	
0.91	0.06	1.18	0.64	18.2	11.7	0.4	19		0.1	
1.37	0.09	2.97	0.01	14.6	1.1	0.2	10		0.0	
1.83	0.10	6.22	0.01	19.5	1.1	0.5	42		0.0	
2.29	0.05	4.88	0.01	19.8	0.2	0.5	41		0.0	
2.74	0.06	1.10	0.62	104	0.8	0.2	15		0.0	
3.20	17.70	0.48	6.66	232	3.3	2.6	1779		0.2	
3.35	25.62	0.51	5.23	254	4.4	7.6	2126		0.7	
3.51	27.20	0.43	3.11	219	7.9	14.7	1170		1.1	
3.66	27.45	0.49	2.77	222	5.0	12.1	954		1.6	
4.11	17.58	1.85	1.63	256	1.7	8.5	348		1.3	
4.57	25.68	1.37	0.69	249	5.9	18.8	861		4.7	
5.03	27.83	1.55	0.76	250	2.7	28.7	717		3.9	
5.49	28.09	2.05	2.06	225	0.5	41.5	366		4.6	
6.10	28.54	1.60	8.76	160	8.0	50.2	306		5.0	
PZ 3										
0.15	0.07	3.36	0.68	11.3	0.5	0.7	74		0.3	50
0.61	0.08	1.76	0.12	18.3	3.0	0.8	9	0.15	0.1	57
1.07	0.19	1.42	0.01	37.8	0.9	1.1	34		0.0	38
1.52	0.46	4.35	0.01	91.8	3.4	1.3	66	0.89	0.1	45
1.98	5.49	0.86	0.26	242	2.1	0.6	691	26	0.0	43
2.13	9.64	0.74	0.86	263	0.5	1.0	1078	43	0.0	62
2.29	15.08	0.57	2.57	264	0.3	19.0	1227	56	0.0	73
2.44	19.48	0.53	4.25	262	31.9	54.4	791	66	0.3	156
2.59	21.42	0.57	5.97	252	1.7	36.0	615	70	0.4	101
2.90	23.47	0.84	6.35	245	0.7	36.1	486	76	1.1	99
3.35	25.40	2.73	4.04	243	0.9	35.5	286	80	5.1	79
3.96	25.65	0.42	2.99	153	0.5	34.5	302	80	11.0	77
4.42	25.94	0.79	4.62	222	39.7	22.8	222		12.5	108
4.57	26.11	0.46	6.88	214	50.6	21.4	182	82	10.3	280
5.03	26.39	0.74	13.1	190	50.6	8.4	156		8.5	98
5.49	26.55	0.74	15.8	161	51.0	4.2	142	84	7.7	99
PZ 4										
0	29.52	6.11	0.82	18.1	202.0	2.0	66		16.5	
0.15	14.87	1.26	1.21	36.4	26.8	3.9	45		2.5	
0.61	7.44	1.66	0.24	76.6	2.6	3.3	16		0.3	
1.07	9.75	1.90	0	55.7	1.6	4.6	22		2.1	
1.68	0.28	2.27	0	47.4	1.9	0.8	13		0.1	
2.13	0.06	4.96	2.68	198	0.7	0.4	12		0.1	
2.59	0.05	1.32	6.06	218	0.6	1.5	22		0.0	
3.05	0.06	1.29	5.04	306	1.7	2.0	39		0.2	
3.51	0.65	1.50	10.24	313	9.7	1.7	6		0.1	
3.96	19.13	1.11	0.91	245	146.0	26.1	258		1.3	
4.42	23.18	0.76	0.48	203	112.0	30.0	171		2.8	
4.88	24.47	0.83	0.71	196	67.4	19.6	171		2.4	
5.33	25.23	0.94	0.36	188	60.4	22.1	150		2.0	
5.79	26.48	0.76	7.8	200	64.5	19.4	150		1.7	
6.25	25.47	1.23	11.71	157	93.8	22.8	157		1.3	
6.71	25.70	0.89	14.14	171	77.6	16.2	151		1.4	
7.16	25.66	1.14	14.15	174	100.9	21.3	456		1.4	
7.62	26.24	1.63	9.57	152	330.0	17.4	181		1.2	
7.92	26.49		8.91	158	491.5	18.9	196		0.8	

Table 1, continued.

Depth (m)	Salinity	Diss. O ₂ (mg L ⁻¹)	pCO ₂ ^a (μmol L ⁻¹)	SiO ₄ ^b (μmol L ⁻¹)	Fe (μmol L ⁻¹)	Mn (μmol L ⁻¹)	Ba (nmol L ⁻¹)	Sr (μmol L ⁻¹)	U (nmol L ⁻¹)	DOC (μmol L ⁻¹)
PZ 5										
0.15	28.84	0.41	9.42	83.4	58.9	14.1	172		4.3	
0.61	28.80	0.39	9.41	121	48.4	23.6	307		3.3	
1.07	25.53	0.78	6.32	244	6.0	37.6	267		1.2	
1.52	25.92	0.57	4.66	260	1.2	43.9	203		1.9	
1.98	25.78	2.92	2.55	189	2.0	9.4	157		8.7	
2.44	25.90	0.40	1.8	200	0.5	8.4	116		10.8	
2.90	25.84	1.38	6.69	175	25.8	3.0	119		4.5	
3.35	25.91	0.76	12.8	159	39.5	3.0	116		4.7	
3.81	26.23	0.77	15.0	150	46.1	3.3	118		4.8	
4.27	26.54	2.34	14.7	142	28.7	3.1	115		5.1	
4.72	26.85	2.46	9.1	132	20.2	2.7	98		3.3	
5.18	27.49	1.28	13.1	141	42.6	2.6	105		3.9	
5.64	28.12	3.80	10.3	138	46.7	2.8	820		2.7	
6.10	28.62	1.14	15.6	121	56.3	1.9	89		2.4	
6.55	28.92	0.52	14.3	102	41.5	1.4	74		2.5	
7.01	28.66	0.46	13.3	401	36.4	1.8	87		2.1	
7.32	28.93	0.37	9.82	371	33.7	2.2	116		1.9	
PZ 6										
0.61	0.08	0.85	2.06	7.93	1.5	0.1	3		2.6	510
1.07	0.23	0.00	8.76	8.72	5.4	0.5	6		1.5	703
1.52	0.35	0.00	0.48	27.4	141.1	5.0	18		0.8	421
1.98	0.10	5.11	2.5	16.9	54.2	1.2	17		1.0	385
2.44	0.06	0.53	4.22	11.3	3.1	5.0	10		0.0	100
2.90	0.07	0.98	5.22	15.6	2.0	2.6	19		0.0	76
3.35	0.08	0.26	5.74	20.8	0.7	0.2	18		0.2	90
3.81	0.08	3.70	4.18	51.5	1.0	0.1	15		0.0	39
4.27	0.14	0.85	8.25	166	2.9	0.1	35		0.1	40
4.72	0.41	0.35	3.8	221	0.2	0.2	33		0.0	33
5.18	6.00	1.42	9.2	354	0.5	37.8	217		0.1	61
5.64	19.64	1.97	2.79	260	0.6	27.4	300		2.4	101
6.10	23.41	0.76	1.28	236	1.2	35.9	250		21.0	174
6.55	24.99	0.65	2.71	211	12.4	42.9	266		20.4	173
7.01	25.74	0.87	1.82	202	42.9	29.1	228		17.3	169
PZ 7										
0.91	0.0	6.40	0.09	245		2.2	10		0.1	300
1.37	0.0	1.08	0.94	181	2.1	0.1	21		0.7	361
1.83	0.0	1.82	1.44	130	10.3	0.7	113		0.8	455
3.20	0.0	1.11	0.74	117	196.3	2.7	22		1.2	510
3.66	0.0	1.87	0.09	109	106.5	12.9	7		0.1	243
4.11	0.0	0.84	0.01	110	100.8	12.0	10		0.2	171
4.11	0.0	0.73	0.01	89.6	67.2	15.7	6		0.4	
4.57	0.0	1.46	0.01	74.6	4.2	8.5	8		0.0	78
5.03	0.0	1.17	0.01	89.0	0.8	0.4	19		0.0	122
5.49	0.0	4.35	0.01	82.7	11.3	0.2	69		0.0	217
5.94	0.0	2.73	0.01	32.4	0.5	0.1	70		0.0	63
6.55	0.0	0.58	2.85	86.9	1.5	0.2	97		0.0	85
7.16	4.00	1.50	1.78	96.6	1.4	56.6	410		0.0	83
PZ 8										
0.15		0.41	5.49	88.1	24.0	3.2	103		7.6	
0.61	27.58	1.99	9.25	178	79.0	31.3	374		3.4	
0.91	28.43	5.06	4.75	232	41.2	30.0	329		1.2	
1.22	25.62	2.78	4.43	278	13.5	42.6	352		0.8	
1.52	25.54	3.52	4.45	272	1.1	43.2	217		0.8	
1.83	25.71	3.02	4.52	240	0.6	33.1	187		1.3	
2.29	25.82	2.73	3.73	221	0.3	19.1	158		2.7	
PZ 9										
0.30	20.04	0.63	1.05	236	1.3	62.3	1893		12.4	
0.61	22.65	1.45	1.8	266	1.8	22.6	591		11.5	
1.07	24.62	2.45	1.54	259	0.9	4.3	532		1.8	
1.52	25.27	0.95	2.49	224	1.2	6.4	389		1.7	
1.83	25.67	4.93	3.12	255	0.2	10.0	331		1.5	
2.29	26.87	1.79	2.6	288	0.3	31.0	244		2.2	
PZ 10										
1.52	0.0	1.12	0.36	8.08	2.9	0.7	73		0.3	
3.05	0.0	0.97	0.24	23.5	119.4	3.0	30		0.3	
3.96	0.0	2.76	0.07	25.1	128.3	4.7	12		1.5	
4.42	0.0	1.37	0.08	2.09	8.6	2.6	8		0.5	
5.49	0.0	5.48	0.06	9.27	1.3	0.2	37		0.3	

Table 2. Surface water trace metal concentrations collected from Waquoit Bay during July 1999. Station locations are shown in Figure 1 of Charette et al. (2001).

Station ID	Salinity	PO_4^{3-} ($\mu\text{mol L}^{-1}$)	SiO_4^{2-} ($\mu\text{mol L}^{-1}$)	Ba (nmol L^{-1})	Sr ($\mu\text{mol L}^{-1}$)	U (nmol L^{-1})
1	31.13	0.81	2.8			11.9
2	31.09	0.81	3.9	45.4		11.6
3	29.92	0.85	10.9	45.4	76.1	10.9
4	29.16	0.78	19.7	56.9		11.1
5	29.32	1.03	24.2	61.3		10.9
6	29.70	1.17	23.6	55.7		11.0
7	30.34	1.20	18.1	66.1		11.1
8	29.71	0.96	17.4	68.8		10.3
9	29.74	0.88	13.8	49.9		10.9
10	30.11	1.00	8.6	52.0		11.0
11	30.99	0.76	3.2	38.8		11.9
12	30.79	0.90	6.9	98.6		11.9
13	31.49	0.77	0.7	22.0	75.8	12.2
14	31.43	1.01	1.3	79.5		12.2
15	31.02	0.72	4.2	39.5		11.7
16	29.61	1.04	20.6	62.0		11.1
17	29.55	0.97	20.9	71.2		11.0
18	30.17	1.10	22.3	90.0		10.9
19	29.75	1.36	11.7	76.3		10.7
20	29.75	1.01	15.6	70.0		11.2
21	31.46	0.71	1.6	34.5		12.4
22	30.36	0.76	15.0	51.2		10.4
23	27.82	0.96	34.9	67.3		10.2
24	27.44	0.90	35.5	64.0		10.0
25	27.61	0.88	34.2	72.6		10.3
26	27.71	0.76	31.4	56.9		9.7
27	27.67	1.02	35.0	78.1	67.7	10.0

Table 3. Major components of the global uranium budget revised to include SGD.

	$\mu\text{mol U m}^{-2} \text{ y}^{-1}$	$10^6 \text{ mol U y}^{-1}$	Reference
<i>Sources</i>			
Rivers		30-60	Palmer and Edmond (1993)
<i>Sinks</i>			
Marine Sediments	0.4	27	Cochran (1982)
Salt Marshes	70	27	Windom et al. (2000)
High/Low Temp. Crustal Alteration		11	Cochran (1982)
Submarine Groundwater Discharge	150	20	This study

Figure Captions

Figure 1. Waquoit Bay on Cape Cod, MA. USA. The mouth at the southern end of the Bay opens to Vineyard Sound. The Bay is approximately 3 km long and 1 km wide and has mean water depth of 1 m. The expanded location map for the head of the Bay shows the piezometer-collected pore water profiles discussed in this paper. Profiles were collected along shore perpendicular (A-A') and shore parallel (B-B') transects in July 2002. PZ9 (not shown) was located 25 m to the east of PZ4.

Figure 2. Two-dimensional contours of pore water salinity (a) dissolved Fe (b) and dissolved Mn (c) in the subterranean estuary at the head of Waquoit Bay. The data were collected from the six piezometer stations along the A-A' shore-perpendicular transect in Figure 1.

Figure 3. Two-dimensional contours of pore water salinity (a) dissolved Ba (b) and dissolved U (c) in the subterranean estuary at the head of Waquoit Bay. The data were collected from the six piezometer stations along the A-A' shore-perpendicular transect in Figure 1.

Figure 4. Pore water profiles of salinity, dissolved Fe, Mn, Ba, Sr, and U for (a) piezometer #3 (PZ3) and (b) piezometer #7 (PZ7) along the A-A' shore-perpendicular transect. PZ7 is located primarily in the freshwater zone of the subterranean estuary. PZ3 covers the whole salinity gradient. The shaded region in (a) represents the salinity transition zone (STZ).

Figure 5. A comparison of pore water and surface water concentrations of dissolved Ba. (a) Pore water concentrations of barium versus salinity for all piezometer samples collected during July 2002. The dashed line indicates the saturation line for barium with respect to barite (barium sulfate) formation (Rushdi et al., 2000). (b) Surface water Ba vs. salinity and the interpolated concentration of Ba at zero salinity.

Figure 6. A comparison of pore water and surface water concentrations of dissolved Sr and Ba for PZ3 collected along the A-A' shore-perpendicular transect. (a) Dissolved Sr vs. salinity for the pore water profile and for three Bay water samples collected in 1999. The solid line indicates

the conservative mixing line between the fresh groundwater and open ocean end members (b) Comparison of the Sr-salinity and Ba-salinity distribution for PZ3. (c) Sr vs. Ba for PZ3; the values 0, 5, 15 and 27 represent the salinity for a given sample.

Figure 7. A comparison of pore water and surface water concentrations of dissolved U. (a) Pore water uranium versus salinity for all piezometer samples collected during July 2002. Also shown is the Bay water U collected during 1999. The dashed line indicates the conservative mixing line for U in seawater (Chen et al., 1986). (b), Expanded U-salinity distribution plot for the surface waters and the interpolated concentration of U at zero salinity.

Figure 8. Barium versus dissolved Mn and barium for the middle three piezometers along the A-A' transect (PZ3, PZ6 and PZ8).

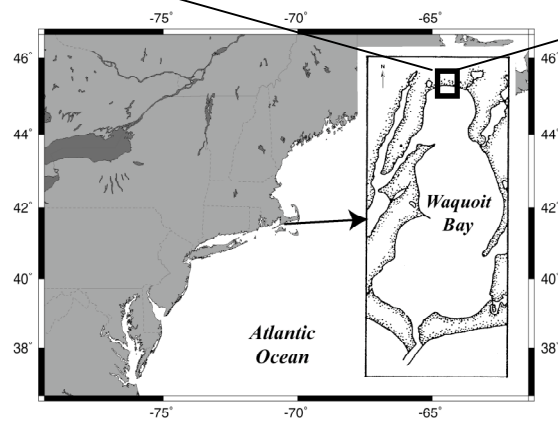
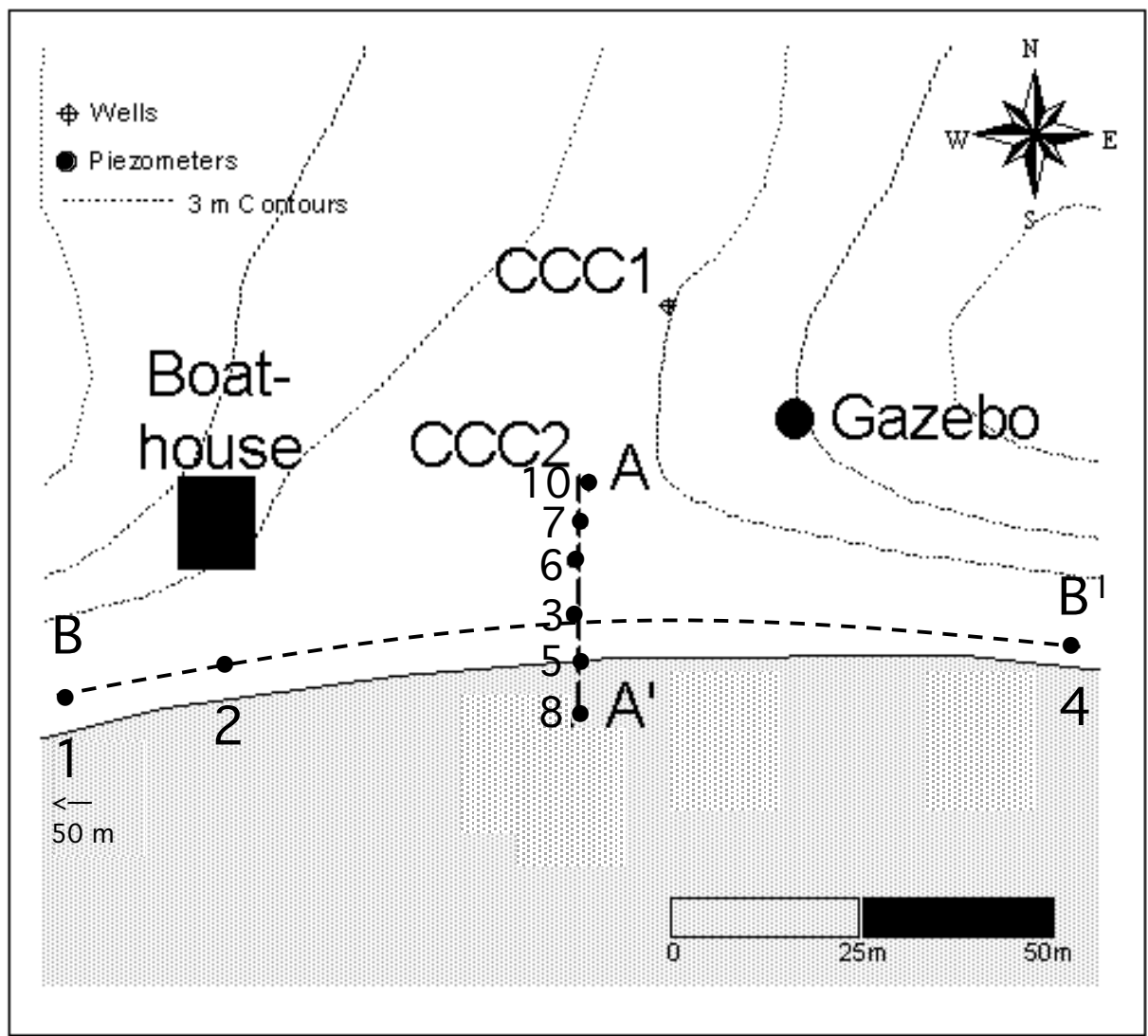


Figure 1

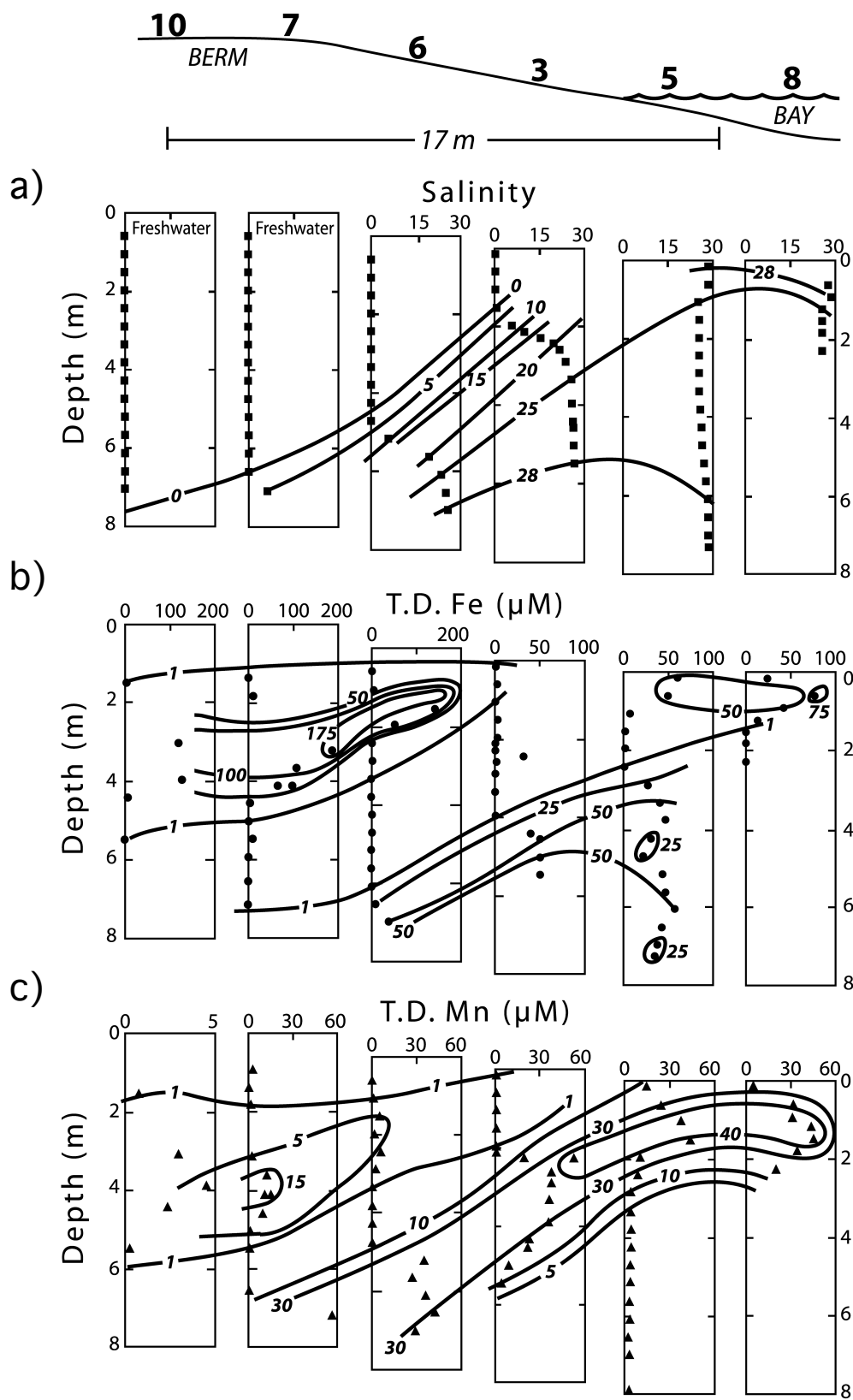


Figure 2

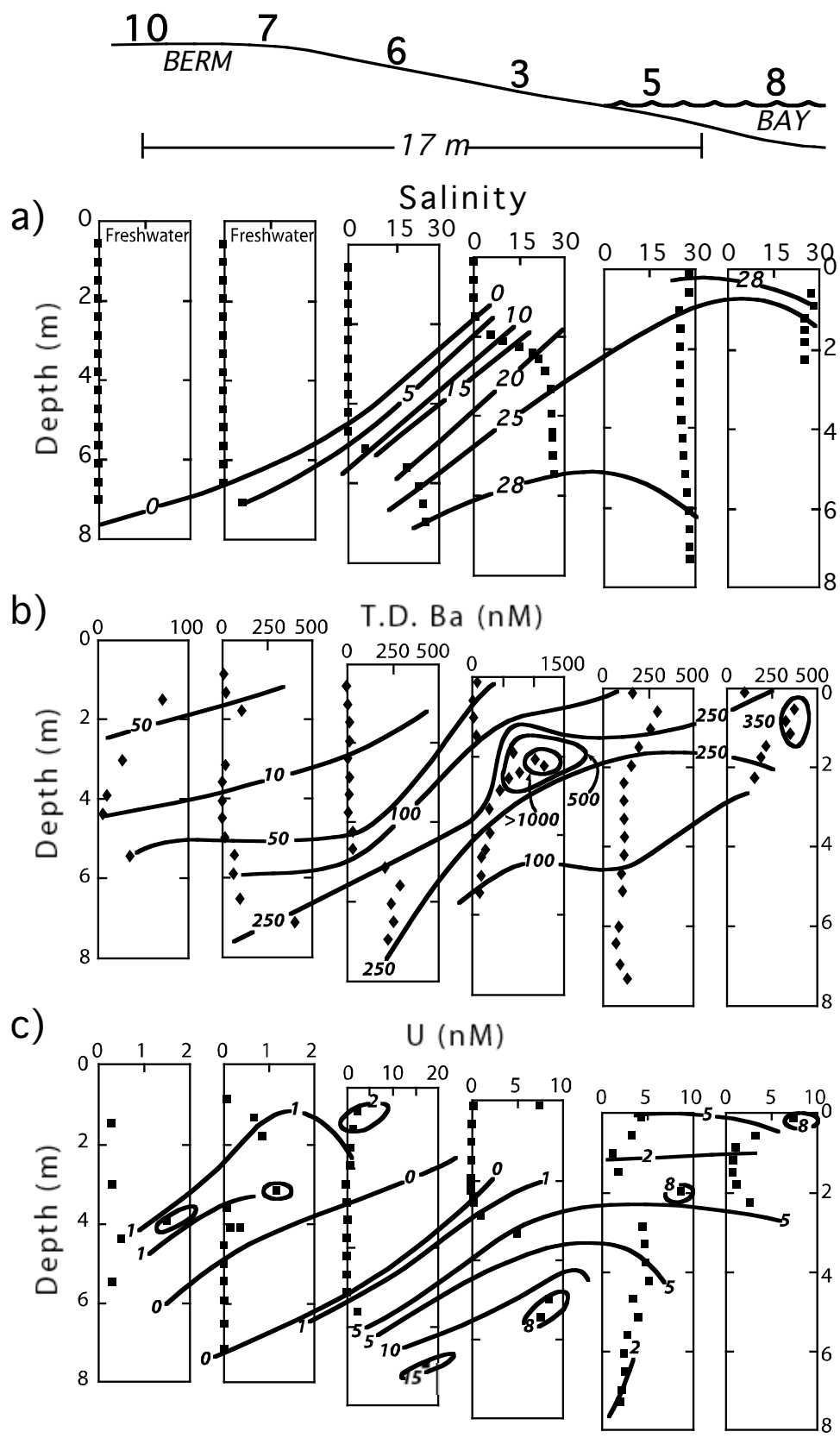


Figure 3

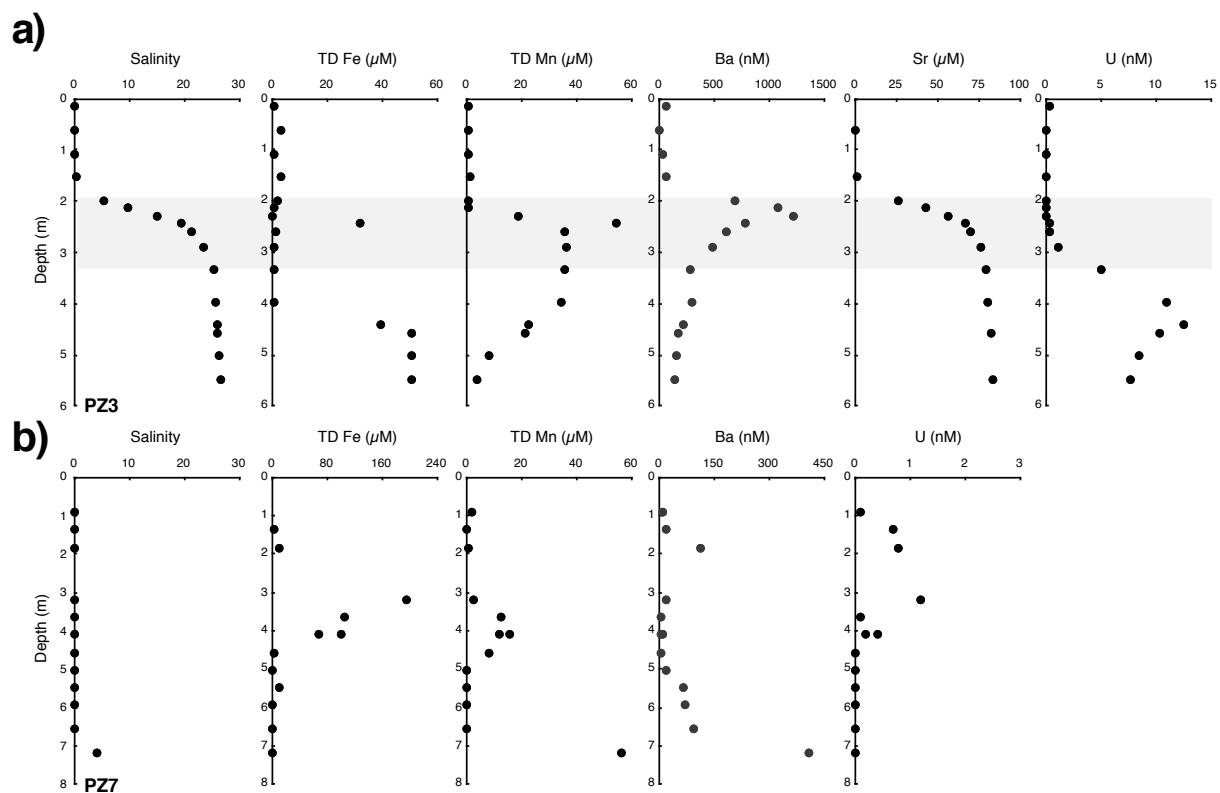


Figure 4

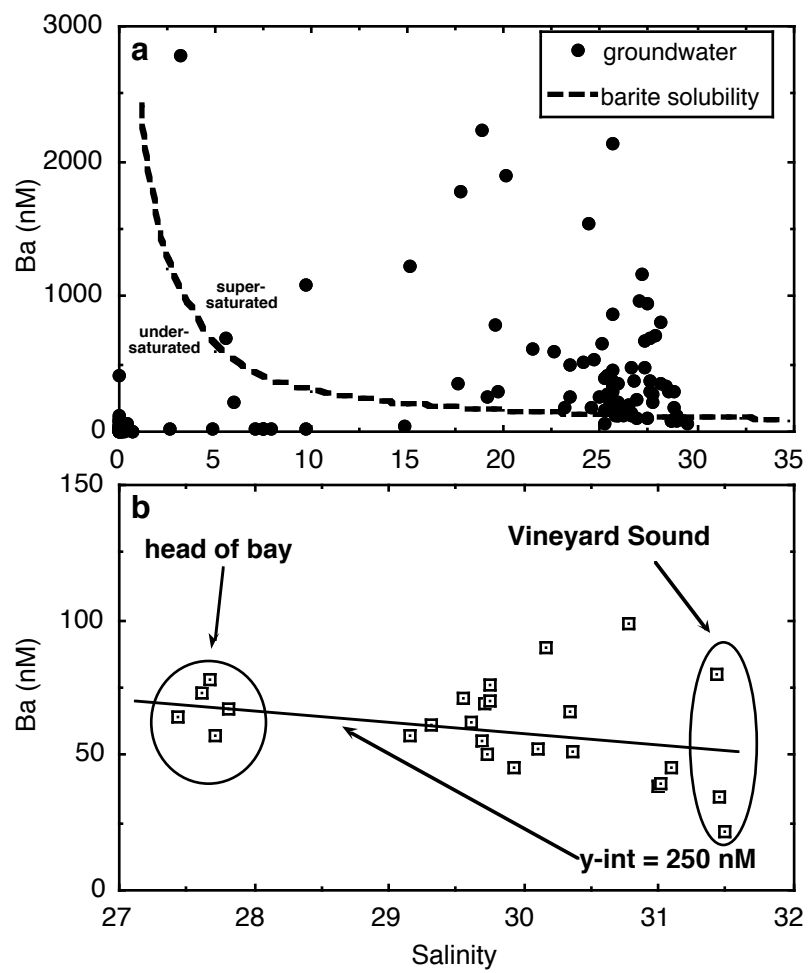


Figure 5

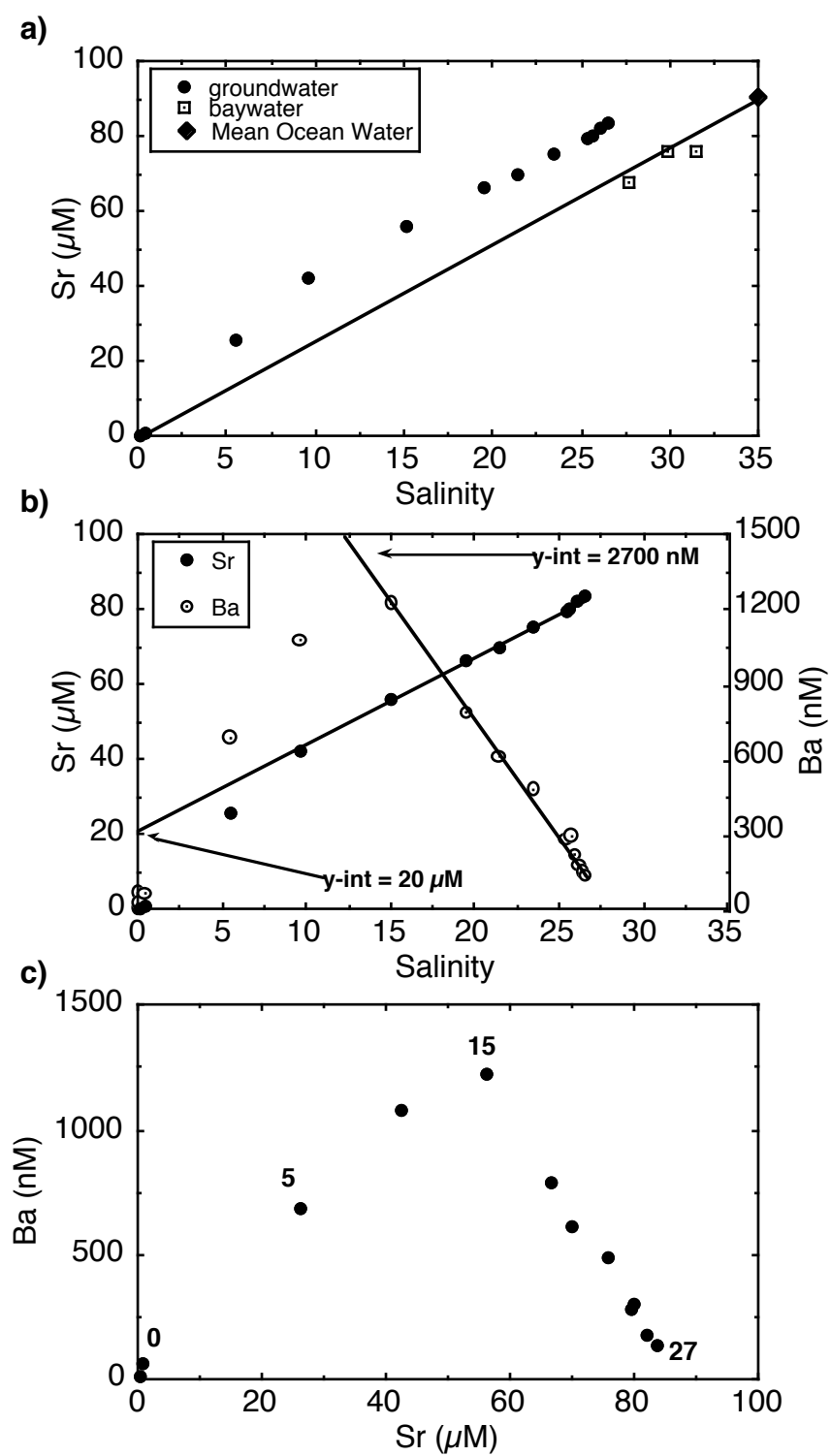


Figure 6

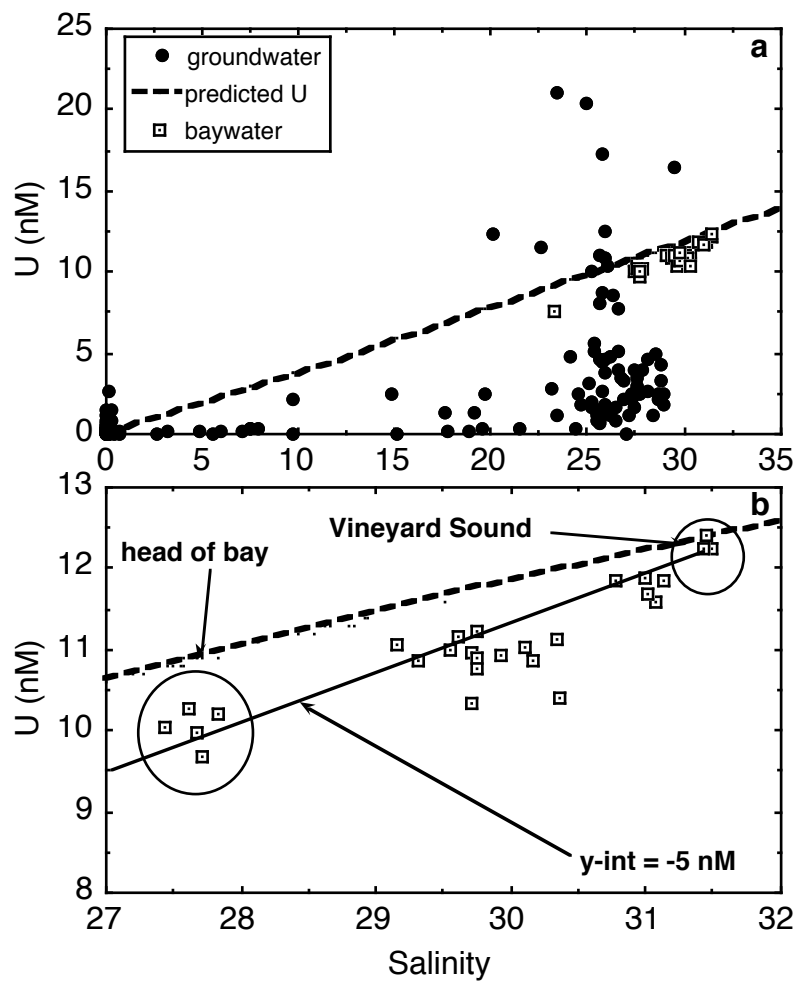


Figure 7

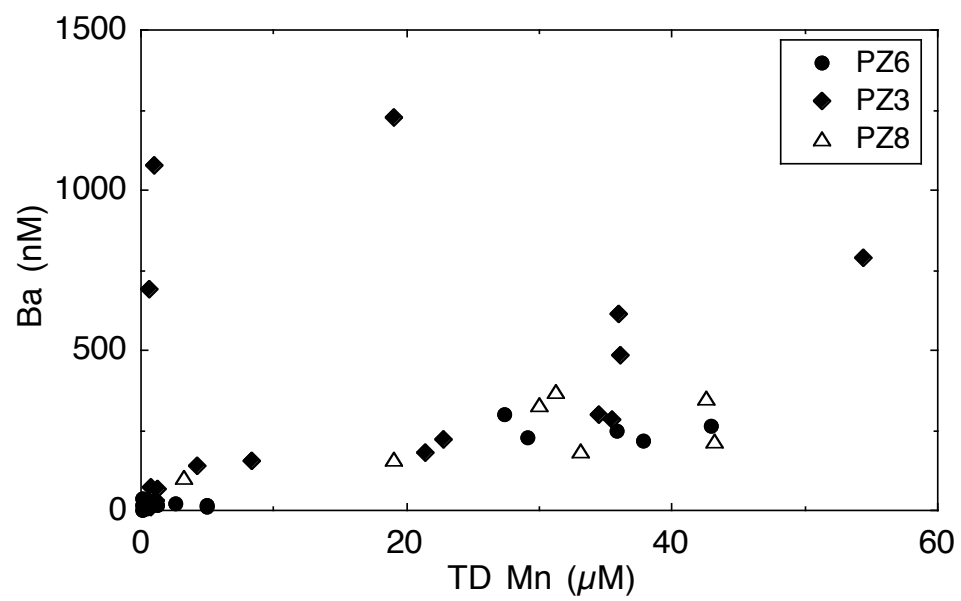


Figure 8